

Page 5 Line 28. It was mentioned that the lower initial concentration of BPA employed higher degradation rate. Once the Y axis value were changed from degradation efficiency (%) to BPA concentration (ppm), the results will be different. How to explain it?

Page 6 Line 6. It is obvious from Figure 7 that ZnO is a more effective catalyst than TiO₂ and SnO₂. The band gap was posed as an explanation. How about other mechanisms? It is better to find more theories from literature as support.

Page 7 Line 10. It is mentioned that ZnO is a better catalyst in photo-oxidation process and it adsorbs more photon-energy than the other photocatalysts. However, as displayed in Figure 10, the advantage of ZnO over TiO₂ and SnO₂ were insignificant. Since no duplicate experiment was carried out, the conclusion might be marginal. If the results were accurate, more explanations about mechanisms need to be discussed further.

It will be better if the results without catalysts can be shown in Figure 7, 9 and 10.

Page 7 Line 18. Electrophilic attack of HOCl on the phenoxide ions was raised as the mechanism of oxidation. In combination with Figure 11, more details can be clarified, e.g. how does the HOCl attack BPA and the degradation compounds step by step; which positions/sites are preferred by HOCl-attacking.

Page 7 Line 31. There was no result which showed ZnO is very stable during the degradation process. Please cite reference articles.